

## Getting Connected HW

Read and outline (Cornell style) **Section 8.3** in your chemistry textbook. Then answer the following assessment questions:

1. What is the difference between a monoatomic ion and a polyatomic ion? Give an example of each.
2. How do you determine the correct subscripts in a chemical formula?
3. Write the correct formula for the ionic compound composed of the following pairs of ions:
  - a. potassium and iodide
  - b. aluminum and bromide
  - c. cesium and nitride
  - d. sodium and nitrate
  - e. magnesium and carbonate
4. Name the following compounds:
  - a. NaBr
  - b. CaCl<sub>2</sub>
  - c. KOH
  - d. Cu(NO<sub>2</sub>)<sub>2</sub> (Hint: uncross-cross to see what charge the transition metal is)
  - e. Ag<sub>2</sub>CrO<sub>4</sub> (Hint: uncross-cross to see what charge the transition metal is)
  - f. CsNO<sub>3</sub>
  - g. K<sub>2</sub>S
  - h. AlF<sub>3</sub>
  - i. CaCO<sub>3</sub>
  - j. Na<sub>3</sub>N
5. Write the formulas for the following ionic compounds:
  - a. Sodium fluoride
  - b. Potassium bromide
  - c. Beryllium chloride
  - d. Iron (III) oxide
  - e. Strontium hydroxide